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Greywater recycling by direct ultrafiltration: understanding fouling

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Abstract

Greywater is no doubt a valuable resource that can be used to alleviate water shortage and increase water conservation in individual households. It is particularly important for arid and semi-arid regions like Australia. Treated greywater can also be used for many activities within the household such as toilet flushing, garden watering, car washing, or pavement cleansing. This study examines the fouling behaviour of submerged ultrafiltration membranes in greywater recycling under concentration variation of common greywater constituents. The results indicate that the use of ultrafiltration directly for greywater recycling in individual households is promising. As expected, increase in particulate matter concentration results in a thicker cake layer. However, hydraulic resistance of such cake layers depends on the presence of other constituents namely calcium and organic matter, which play a major role in membrane fouling. Fouling increases linearly with organic matter concentration, while the presence of a small amount of calcium may enhance fouling significantly.

Keywords: Ultrafiltration, greywater, fouling, calcium, particulate matter, organic matter, water recycling

1. Introduction

Widespread water scarcity and the growing stress on water supply resources have prompted considerable interest in water recycling together with other water saving measures. The emphasis is, however, mostly on domestic wastewater and on large-scale decentralised options [1]. Although, there have been a number of studies on greywater recycling focusing on large complexes such as commercial buildings [2] or sport stadiums [3], its potential has not been adequately investigated. This is particularly true for decentralised greywater recycling systems for individual households.

Greywater consists of all non-toilet wastewater generated in a household. It includes wastewater from showers, baths and spas, hand basins, washing machines, laundry troughs, dishwashers, and kitchen sinks, although kitchen sink wastewater is often excluded from recycling due to the high organic and microbial load. In Australia, greywater represents most of in-door water consumption and is sufficient for other water intensive activities that do not necessarily require potable water standard such as toilet flushing and garden watering [4].

Greywater contains much less pathogens than general domestic wastewater, although the concentration of other constituents can vary significantly as their origins are activity-specific. Consequently, greywater treatment systems must be robust enough to sustain from contaminant shock loads. Ultrafiltration membranes can produce consistent and reliable water quality. However, its fouling behaviour in response to such variation in greywater composition and constituent concentrations has to date not been adequately studied. The objectives of this study are to evaluate the feasibility of submerged ultrafiltration in greywater recycling under realistic concentration variations of common greywater constituents and to understand the impact of those constituents on fouling behaviour.

2. Greywater characterisation

The characteristics of greywater from a single household depend mostly on the activities in the household, the quality of the water supply, and to a lesser extent, the type of distribution system for both potable water supply and greywater collection. Consequently, characteristics may vary significantly from source to source. This has important implications on the evaluation of treatment technology and subsequent reuse possibilities. An in-depth literature review was conducted to determine major greywater constituents and their concentration ranges. The results are presented in Table 1, in comparison with general domestic wastewater.

Table 1 Composition of domestic wastewater and greywater [2, 5-9]

Parametre	Domestic wastewater	Greywater	Unit
pH	5.9 - 7.7	5 -10.9	(-)
Total solids	200 - 1200	113 - 2410	(mg/L)
Total dissolved solids (TDS)	100 - 850	137 - 1260	(mg/L)
Suspended solids (SS)	24 - 608	3.1 - 330	(mg/L)
BOD5	28 - 400	33 - 1460	(mg/L)
TOC	42 - 290	30 - 880	(mg/L)
COD	45 - 1000	3.8 - 1380	(mg/L)
Nitrogen	20 - 117	0.28 - 74	(mg total N /L)
Phosphorus	3.2 - 30	0.0062 - 74	(mg total P /L)
Calcium	1.1 – 200	3.6 - 200	(mg Ca /L)

A review of published literature reveals that greywater constituents are similar to those of domestic wastewater. However, greywater usually contains much less pathogens, hence, is much more suitable (and socially more acceptable) for recycling options in a single household context. Concentration ranges of most common constituents are considerably wider than that of wastewater. This implies that treatment systems must be robust enough to sustain peak loads and variation in influent quality.

3. Materials and Methods

3.1. Membrane module

A ZeeWeed®-1 (ZW-1) submerged hollow fibre ultrafiltration module supplied by Zenon Environmental (Canada) was used in this study. This module operates in an "outside-in" configuration and has a nominal pore diameter of 0.04 µm, with an effective membrane surface area of 0.047 m². The membrane has a pure water flux of 64 L.m².h⁻¹ at a transmembrane pressure of negative 15 kPa, corresponding to a clean membrane resistance of 9.6 ·10⁻¹¹ m⁻¹. According to the manufacturer, the membrane is chlorine resistant.

3.2. Filtration set up

The filtration set up consisted of a membrane module, a peristaltic pump, and a pressure gauge. The membrane module was placed in a 2.25 L reservoir. The feed solution was constantly agitated using a magnetic stirrer. The system was equipped with a fish tank aerator for membrane cleaning purposes. This system is described in detail elsewhere [10].

3.3. Analytical

UV_{254nm} absorption was measured using a Shimadzu UV-Visible Spectrophotometer as an indicator for HA concentration in the absence of particulates. Particle size distributions of cellulose and kaolin were measured using a Mastersizer (Malvern Instrument, UK). A Hach 2100N turbidity meter was used for turbidity measurement. Total organic carbon (TOC) was analysed using a Shimadzu TOC-S analyser. Calcium was analysed using a Varian Atomic Absorption Spectrometer.

A "Micro 2 pH/Ion Meter" was used to determine pH. TOC, UV_{254nm}, turbidity, and pH measurements were undertaken immediately after sampling, while samples were stored at 4°C for later calcium analysis.

3.4. Chemicals

Kaolin, cellulose, technical humic acids were purchased from Sigma Aldrich. Home Brand household bleach was purchased from a local department store. This bleach contained 3.5% sodium hypochlorite as the active ingredient. All other chemicals were of analytical grade.

As a result of the literature review carried out for this study, key greywater constituents were selected and used to prepare synthetic greywater for experimental design (see Table 2). The concentration of these constituents was then varied within the range commonly found to examine the fouling phenomena during ultrafiltration of greywater for recycling purposes.

Table 2 Synthetic greywater constituents chosen for this study

Constituent	Particle Size*	Representative for component/pollutant group
NaCl	n/a	Dissolved monovalent salts
NaHCO ₃	n/a	Natural Buffer
Kaolin	D ₁₀ = 0.54 µm D ₉₀ = 9.6 µm	Suspended inorganic matter, inorganic colloidal particles
Cellulose	D ₁₀ = 2.95 µm D ₉₀ = 49.1 µm	Suspended organic matter, organic colloidal particles
Humic acid	n/a	Dissolved organic matter
CaCl ₂	n/a	Dissolved multivalent salts

* Determined in this study; D₁₀ (or D₉₀) value presents the diameter at which 10 % (or 90 %) of the particulate matter diameter is smaller than this value.

n/a: Not applicable or not available

3.5. Experimental protocol

The synthetic greywater solution was prepared 16–18 h before the experiment and stored at 4°C to facilitate equilibration. All experiments were conducted with a constant flow rate of 50 mL.min⁻¹ for a duration of 6 h. Aeration was not applied during the experiment to allow fouling development. Permeate was recirculated back to the feed reservoir. The transmembrane pressure (TMP) was recorded every 30 min to calculate flux and hence measure fouling. Feed and permeate samples were also collected for analysis at the same interval. Unless otherwise stated, synthetic greywater used in this study contained concentrations as follows: humic acid 20 mg/L, kaolin 50 mg/L, cellulose 50 mg/L, CaCl₂ 0.5 mM (corresponding to 20 mg/L of Ca), NaCl 10 mM, and NaHCO₃ 1mM at pH 8.

At the end of each experiment, the membrane was aerated for 15 min to loosen and partially remove the cake layer. The membrane was then soaked and aerated in a 0.5 mg/L surfactant (sodium dodecyl sulphate – major ingredient in shampoo) solution, followed by bleach cleaning. The bleach concentration of the cleaning solution was 1 % by volume. Both of these chemicals are commonly used within the household. Pure water flux and turbidity retention were checked at the end of the cleaning process.

Parameters used to quantify the efficiency of UF, are the increase in TMP (Δp) as an indicator for fouling, where TMP₀ is the transmembrane pressure at the start of the experiment and TMP_E is the transmembrane pressure at the end of the experiments

$$\Delta p = \text{TMP}_E - \text{TMP}_0 \quad (1)$$

and solute rejection (R) is calculated as indicated in Eq (2), where c_F and c_P are feed and permeate concentrations at the time of measurement, respectively.

$$R = 100 \cdot \left(1 - \frac{c_P}{c_F} \right) \quad (2)$$

4. Results and Discussion

4.1. Membrane fouling and cake formation

Cake formation on the membrane surface was visible at the end of all experiments. Snapshots of the feed container and membrane at different times of a typical experiment are shown in Figure 1. Interestingly, fouling (or increase in TMP) was only recorded under certain conditions, despite the formation of a cake layer. For all experiments, turbidity of all permeate samples was less than 0.4 NTU, despite of a very high turbidity (140 NTU at the beginning of the experiment) in the feed solution. This will be further elaborated in later sections.

Preliminary experimental results indicate that if intermittent aeration is applied, fouling will take significantly longer to develop, which is consistent with results reported by other researchers [11-13].

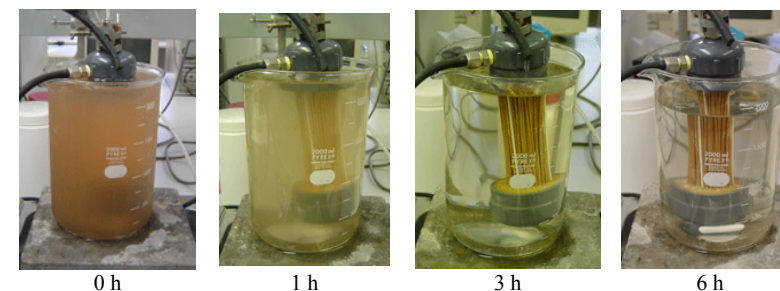


Figure 1 Feed solution turbidity and cake deposition as a function of filtration time (pH 7.5-8.0, HA 20 mg/L, Kaolin 50 mg/L, Cellulose 50 mg/L, CaCl₂ 3 mM, NaCl 10mM, NaHCO₃ 1mM).

Following the cleaning protocol as outlined above, no discernible variations in membrane permeability and turbidity retention of the module were observed. This indicates that common household chemicals (bleach) may be effectively used for membrane cleaning which facilitates the maintenance of decentralised systems, although findings are yet to be trialled in long term tests.

4.2. Calcium effects on fouling & retention

Calcium has been recognised as an important multivalent cation contributing to membrane fouling through the complexation with organic matter and possibly with membrane functional groups [14].

As can be seen in

Figure 2A, calcium can significantly influence membrane fouling during ultrafiltration of greywater also. While there is some fouling in the absence of calcium (ΔP of 4 kPa) and no fouling at all at high calcium concentration (>3 mM), a peak of membrane fouling was found at low calcium concentration in the vicinity of 0.5 mM (ΔP of up to 12 kPa).

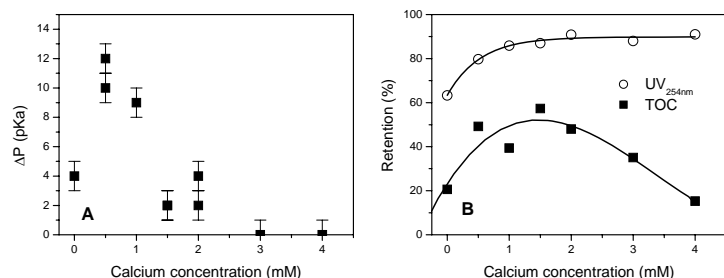


Figure 2 Calcium concentration effects on fouling (A) and retention (B) (pH 8, HA 20mg/L, kaolin 50 mg/L and cellulose 50 mg/L, NaCl 10 mM, NaHCO₃ 1mM). ΔP is the transmembrane pressure increase due to fouling as described in eq (1).

While a similar phenomenon has also been reported for nanofiltration [15] and microfiltration [16] membranes for water and wastewater treatment applications, the actual mechanism remains unclear. Aoustin *et al.* [17] studied the fouling of two ultrafiltration membranes due to natural organic matter and reported negligible calcium concentration effects. However, in their study, colloidal particles were not present in the feed solution. It is well known that calcium can act as a coagulant via charge neutralisation and complexation-bridging between macro-organics [14, 18]. It is hypothesised that high calcium concentration may result in the formation of large and compact aggregates (or flocs) that are bigger than the pores and hence cannot penetrate into the membrane and cause pore blocking. Further, the cake layer formed under such conditions may be more porous and possesses a negligible hydraulic resistance. It is known that colloidal systems at lower calcium concentrations form loose and smaller aggregates that result in a dense cake layer or pore blocking. This is consistent with the results reported by Waite *et al.* [18] who found a much larger fraction of small aggregates under low ionic strength compared to high ionic strength where most of aggregates were much larger. This hypothesis is also supported by the fact that TOC retention as a function of calcium seems to somewhat mirror fouling (see

Figure 2B); highest TOC retention coincides with the highest level of fouling. TOC presents all dissolved organics in the solution while UV_{254 nm} absorption tends to be an indicator of larger aromatic fractions of the organics but also measures particles by light scattering effects rather than absorption. As colloidal particles are almost completely retained by the membrane in all cases, UV_{254nm} retention is significantly higher than that of TOC. This “interference” of colloidal and particulate matter also explains the difference between TOC and UV_{254 nm} retention curves as a function of calcium concentration.

The role of calcium as an important fouling ingredient is further evident in Figure 3, which shows a decrease in calcium concentration in both permeate and feed solutions over time. This clearly indicates the complexation of calcium and organic matter and subsequent deposition in the fouling layer.

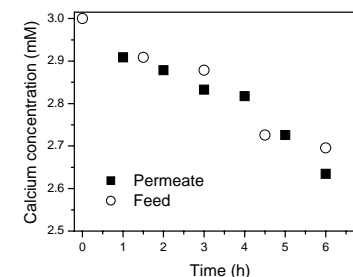


Figure 3 Calcium concentration of feed and permeate as a function of time (pH 7.5-8.0, humic acid 20mg/L, kaolin 50 mg/L, cellulose 50 mg/L, CaCl₂ 3 mM, NaCl 10 mM, NaHCO₃ 1mM).

4.3. Effect of particulate matter type and concentration on fouling & retention

Particulate matter is arguably ubiquitous in greywater. Size characterisation measurement (data not shown) indicates that both cellulose and kaolin used in this study are quite large as compared to the membrane pore. D₁₀ value of cellulose and kaolin are 2.95 μ m and 0.54 μ m, respectively (see Table 2). In comparison, the membrane pore diameter is substantially smaller at 0.04 μ m. This implies that the cake layer formed by the accumulation of particulate matter on the membrane surface can be very porous and may not result in measurable fouling if no other constituents are present.

Effects of cellulose and kaolin concentration on fouling are shown in

Figure 4. Fouling increases marginally as cellulose concentration increases, whereas an increase of kaolin concentration beyond 100 mg/L does not result in further fouling. The effects of particulate matter concentration on TOC and UV_{254 nm} retention appear to be negligible (data not shown).

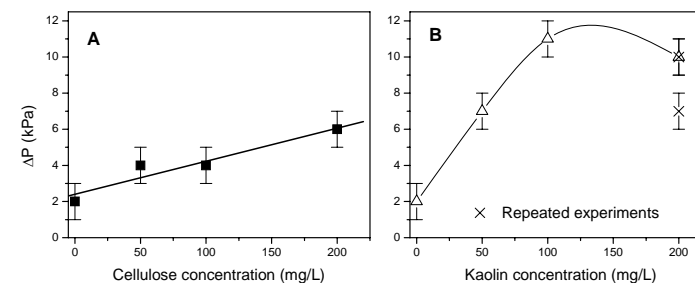


Figure 4 Cellulose (A) and kaolin (B) concentration effects on membrane fouling (pH 7.8-8.9, Ca²⁺ 0.5 mM, HA 20mg/L, NaCl 10mM, NaHCO₃ 1mM).

4.4. Effect of organic matter concentration on fouling & retention

Organic matter found in greywater can be much smaller than the pore size of the membrane, and can cause significant fouling as can be seen in Figure 5A. Humic acid (HA) concentration correlates strongly with membrane fouling. The transmembrane pressure (presented by ΔP) increases rapidly with increasing HA concentration, presumably due to adsorption to particulate matter in the cake layer, which result in a more compact and denser cake layer. This is further enhanced by the presence of a small amount of calcium (0.5 mM) which acts as a bridging ion and

facilitates HA absorption to particulate matter and the membrane surface. HA is also known to foul membranes by gel formation and adsorption in the membrane pores [15].

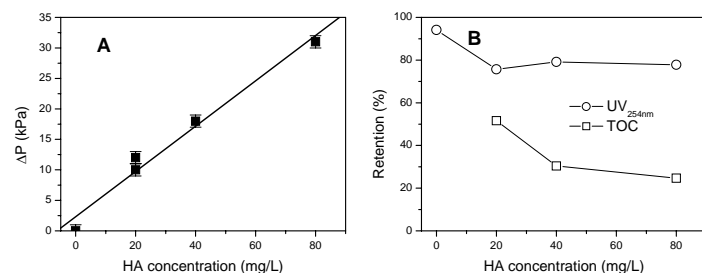


Figure 5

HA concentration effects on fouling and retention (pH 7.6-8.3, Ca²⁺ 0.5 mM, Kaolin 50 mg/L, Cellulose 50mg/L, NaCl 10mM, NaHCO₃ 1mM).

High concentration of HAs within the cake layer can also result in a significant decrease in TOC retention (reduction from about 50 to 20%) as HA concentration increases (see Figure 5B). However, HA concentration appears to have negligible effect on UV_{254 nm} retention, which can be explained with the high retention of particulate matter. It is expected that large organic compounds (which contribute to UV_{254 nm} adsorption) have a stronger tendency to cause pore blockage in association with organic and inorganic colloids as compared to small organics [19]. These small organics preferentially permeate through the membrane, resulting in a lower TOC retention, which would be further enhanced by stronger concentration polarisation effects at the higher HA concentration.

5. Conclusions

The results reported in this study indicate that common household chemicals (bleach) may be effectively used for membrane cleaning which facilitates the maintenance of decentralised systems, although findings are yet to be trialled in long term tests.

Calcium and organic matter play a major role in membrane fouling. Fouling increases linearly with organic matter concentration. While the presence of a small amount of calcium enhances fouling significantly, fouling may be completely diminished at calcium concentrations larger than 3 mM.

Increase in particulate matter concentration results in a thicker cake layer. However, hydraulic resistance of such cake layers depends on the presence of other greywater constituents such as calcium and dissolved organic matter.

Further work is in progress to evaluate the impact of other foulants such as proteins and surfactants, effects of foulants on retention of trace contaminants and the successful continuous operation and maintenance of larger scale units.

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